

12. Nitrogen and sulfur

12.1 Nitrogen and sulfur

Paper 2

Marking Scheme

Q1.

(a)(i)	(natural =)lightning	1
	(man-made =)internal combustion engines	1
(a)(ii)	$2\text{NO}_2 + \text{H}_2\text{O} \rightarrow \text{HNO}_2 + \text{HNO}_3$ OR $4\text{NO}_2 + \text{O}_2 + 2\text{H}_2\text{O} \rightarrow 4\text{HNO}_3$	1
(a)(iii)	It / NO_2 reacts with (unburned) hydrocarbons / VOCs ALLOW reaction of unburned hydrocarbons / VOCs in presence of NO_2	1
(a)(iv)	$2\text{HNO}_3 + \text{CaO} \rightarrow \text{Ca}(\text{NO}_3)_2 + \text{H}_2\text{O}$	1
(a)(v)	brown fumes given off	1

Q2.

(b)(i)	M1 • high temperature • high pressure • N_2 Any two correct for M1 M2 $\text{N}_2 + \text{O}_2 \rightarrow 2\text{NO}$ OR $\text{N}_2 + 2\text{O}_2 \rightarrow 2\text{NO}_2$ OR $\text{N}_2 + x\text{O}_2 \rightarrow 2\text{NO}_x$	2
(b)(ii)	M1 NO_2 is the catalyst (for $\text{SO}_2 \rightarrow \text{SO}_3$) OR catalytic oxidation (of SO_2) by NO_2 M2 $\text{NO}_2 + \text{SO}_2 \rightarrow \text{NO} + \text{SO}_3$ AND $2\text{NO} + \text{O}_2 \rightarrow 2\text{NO}_2$ M3 SO_3 reacts with rain OR moisture/water (in atmosphere) to form sulfuric acid / H_2SO_4 OR $\text{SO}_3 + \text{H}_2\text{O} \rightarrow \text{H}_2\text{SO}_4$	3
(b)(iii)	photochemical smog / PAN	1

Q3.

(d)	<i>natural process</i> : lightning	1
	<i>man-made process</i> : internal combustion engines	1
(e)(i)	species with 1 or more unpaired electrons	1
(e)(ii)	$\text{NO}_2 + \text{SO}_2 \rightarrow \text{SO}_3 + \text{NO}$	1
	$\text{NO} + \frac{1}{2} \text{O}_2 \rightarrow \text{NO}_2$	1
(e)(iii)	(formation of) acid rain	1

Q4.

(a)(i)	$4\text{P} / \text{P}_4 + 5\text{O}_2 \rightarrow \text{P}_4\text{O}_{10}$	1
	$\text{P}_4\text{O}_{10} + 6\text{H}_2\text{O} \rightarrow 4\text{H}_3\text{PO}_4$	1

Q5.

(b)(i)	M1 react with (unburnt) hydrocarbons M2 (form) PAN / peroxyac(et)yl nitrate	2
(b)(ii)	$2\text{NO} + 2\text{CO} \rightarrow 2\text{CO}_2 + \text{N}_2$ OR $\text{NO}_2 + 2\text{CO} \rightarrow \frac{1}{2}\text{N}_2 + 2\text{CO}_2$	1
(c)	any Group 1 hydroxide or $\text{Ca}(\text{OH})_2$ / $\text{Sr}(\text{OH})_2$ / $\text{Ba}(\text{OH})_2$	1

Q6.

(d)	lightning	1
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Q7.

(a)	$4\text{NH}_3 + 5\text{O}_2 \rightarrow 4\text{NO} + 6\text{H}_2\text{O}$	1
(d)	M1 state the effect of NO gas on contact with moist air NO reacts with water OR NO reacts with oxygen and water. M2 consequence of M1 in terms of atmospheric pollution causing acid rain OR photochemical smog / ground level ozone OR destroy ozone layer	2
(f)	fertiliser	1